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(pp. 280-282) VOCABULARY Complete each
sentence. 1. ... Circle the letter of the
best answer. 1. Given the balanced equation

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$2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$, how many moles of CO_2 are produced when 14.9 g of O_2 are used? a. 0.29 mol b. 3.49 mol

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Reactions Chapter Review F f. Pentane burns
in oxygen to produce carbon dioxide and water
 $C_5H_{12} + 8O_2 \rightarrow 5CO_2 + 6H_2O$ 47. Classify the
following reactions as single displacement,
double displacement, Page 9/26

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