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Consists of a
single atom with
a positive or
negative charge
resulting from

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the loss of gain
of one or more
valence
electrons and
behaves as a
unit, yet it
still has a
charge. (+ or
-). Determined
by the amount of
electrons lost
or gained.

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To name a binary molecular compound, use the following guidelines: 1. Write the names of the elements in the order listed in the formula 2. Use

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prefixes

appropriately to indicate the number of each kind of atom 3.

End the name of the second element with the suffix -ide

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First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix

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-mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the anion name

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ends in -ite the
acid name is the
stem of the
anion with the
suffix -ous
followed by the
word acid (ex.
H₂SO₃ --
X=sulfite =
sulfurous acid)

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Writing the
Names and
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Bases. A base is
a compound that
produces OH^- in
water. When
naming a base,
you name it like
any other
compound that
starts with a
regular or

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transition

metal. Ex: NaOH

= sodium

hydroxide. When

writing the

formula for a

base, remember

to balance

charges. Ex:

magnesium

hydroxide =

Mg(OH) 2

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NAMING AND

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FOR MOLECULAR

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COMPOUNDS (pages
268–270) This
section explains
the rules for
naming and
writing formulas
for binary
molecular
compounds.

Naming Binary
Molecular

Compounds (pages
268–269) 1.

Circle the

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letter of the
type(s) of
elements that
form binary
molecular
compounds.

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~~SECTION 9.2~~

~~NAMING AND~~

~~WRITING FORMULAS~~

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COMPOUNDS I.

Write the formulas for these binary ionic compounds.

c. potassium iodide
2. write the formulas for the compounds formed c. b. f.

3. Name the following binary ionic compounds.
sodium sulfide

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g. $\text{CuCl}_2 \cdot \text{h.}$

snc. l. a. MnO_2

c. d. SrBr_2 221

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a. 2- b. 1+ c.

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1- d. $3+$ 43. a.

$2+$ b. $2+$ c. $3+$

d. $1+$ 44. a.

barium ion b.

iodide ion c.

silver ion d.

mercury(II) ion

45. cyanide, CN^-

and hydroxide,

OH^- 46. a.

hydroxide ion b.

lead(IV) ion c.

sulfate ion d.

oxide ion 47.

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zero; A compound is electrically neutral. 48. The symbols for the cation and

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help yourself to
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vocabulary and
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presented in
this chapter. A
B; anion: any

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atom or group of
atoms that has a
negative charge:
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Identify the charges of monatomic ions by using the periodic table, and name the ions. STA:

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